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| Review unit 1-Ch1.1, Ch2,Ch3.3,Ch4,Ch5.1, 5.2, Ch 6.1,6.2 Chem 11-Mrs. Sanford  Name: period: |
| Matter: |
| Chemistry: |
| Five areas of Chemistry: |
| Extensive property |
| Intensive property |
| Physical property |
| States of matter 1- 2- |
| 3- |
| Physical change: Give 3 ex:1- 2- 3- |
| Two types of mixtures and a description for each:  1-  2- |
| Solution: Phase: |
| Ways to separate matter: |
| 1. 2-   3 4  5 6  7 8 |
| Element: compound: |
| Chemical change: |
| A substance differs from a mixture in that the composition: |
| Chemical property: |
| Four things that tell us a chemical change has taken place:   1. 2. 3. 4. |
| Precipitate: |
| Law of Conservation Of Mass: |
| Atomic Number- Mass Number- |
| Number of neutrons= |
| Isotope: Give an example of two isotopes of Neon 🡪 |

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| Scientist | Key points of their discovery |
| Democritus |  |
| Dalton |  |
| JJ Thomson |  |
| Nagaoka |  |
| Rutherford |  |
| Bohr |  |
| deBroglie |  |
| Shroedinger/Heisenberg |  |
| Chadwick |  |
| Einstein |  |
| Mendeleev |  |

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| Diagonal rule: Draw the energy level filling diagram :  AND Give the electron configuration for sulfur and barium  2s  1s |
| An s orbital can hold a maximum of\_\_\_\_\_\_\_\_electrons....a p orbital can hold a maximum of\_\_\_\_\_\_electrons....d orbital\_\_\_\_\_\_electrons and f orbital\_\_\_\_\_\_electrons |

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| family | Group or location |
| Alkali metals |  |
| Alkaline earth metals |  |
| Halogens |  |
| Metals |  |
| Non-metals |  |
| Transition metals |  |
| Lanthanides |  |
| Actinides |  |
| metalloids |  |
| Metric conversion: Draw table of conversion prefixes on the reverse of this page.  Do the following questions p 84 32 all,33b to c, p 87 42 all, 44 all, p 153 #1, p 111 #15, 16, p 112 #17 and p 113 #19 | |